

Penny Ante Equilibrium Lab Answers



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Penny Ante Equilibrium Lab Answers

Lab - PENNY-ANTE EQUILIBRIUM Equilibrium is a state where there are no observable changes over time. We have studied examples of physical equilibrium, such as the process $\text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_2\text{O}(\text{g})$. Equilibrium is achieved when the rate of the forward reaction equals the rate of the reverse reaction.

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Penny Equilibrium Lab Activity ... Write a definition of equilibrium based on your answer to Question #1. 3. Compare the results obtained in Parts A and B. A. ... The P/R ratio may be called the "equilibrium constant" for the penny reaction. What does that mean? B. Does the container the pennies start in affect the final ratio of the ...

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1 Name(s): Date: LAB PENNY-ANTE EQUILIBRIUM Equilibrium is a state where there are no observable changes over time. We have studied examples of physical Equilibrium— Experiment Summaries and Concepts

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Summer Intro Letter; Summer Assignment (partial KEY) This is the review assignment you need to complete.It will be collected within the first week of the semester.Take the time to make sure can do more than just get an answer on paper - you need to understand how you got there!

AP Documents - Class Website ~ Mrs. Riggs

The zinc (which was already silver) coated the penny. The NaOH dissolved the zinc. (that was what made the water murky) And small pieces of zinc adhered to the penny (coated) It was a physical change because no NEW color was created.

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3. The ratios do not illustrate the concept of law of mass because law of mass is describing the reaction at a given temperature which in this case, it did not include temperature. It does however illustrate equilibrium because of the reversible reactants. 155 Conclusion 40 20 15

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Title Le Chatelier's Principle Demos: Description This is the full instructions for two demos, both about the concept of reversible reactions and introducing Le Chatelier's Principle.

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Penny-Ante Equilibrium Activity. Introduction: What is equilibrium? What happens to the amount of reactants and products when equilibrium is reached? What if more reactants or products are added to a system already at equilibrium? ... Write a definition of equilibrium based on the answer to Question #1. Compare the results obtained in Parts A ...

Name _____ **Date** ...

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